

Correction to Three-Dimensional Homogeneous Ferrite-Carbon Aerogel: One-Pot Fabrication and Enhanced Electro-Fenton Reactivity

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“Under low pH conditions, the process appears to be controlled by redox cycling of surface and dissolved iron (i.e., Fe[II]/Fe[III]), the latter resulting from dissolution of iron oxides.^{48,52,53} At neutral pH value, the contribution of dissolved iron to H₂O₂ activation should be minimal because Fe(III) is indissoluble.^{54,55} Therefore, the decomposition of H₂O₂ under neutral pH condition is considered as a surface-catalyzed process.”

“... presumably fits the Haber–Weiss mechanism that is analogous to that observed in the homogeneous Fenton system.”

“.....not catalyze the decomposition of H₂O₂. Thus, the significant enhancement in H₂O₂ activation.”

“The role of steric position of reactive sites on redox processes has been deduced to be important in the reduction of carbon tetrachloride by Fe(II) associated with goethite, where the steric position of the latter can enhance multiple electron transfer reactions.^{48,61} In a similar way, iron dispersion within the carbon matrix might favor those radical mechanisms above-mentioned, leading to more •OH production during the decomposition of H₂O₂.”

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